

## Chemistry-Unit 2 Test

1. Any change not involving a change in the substances chemical make-up.
  - a. Chemical change
  - b. Physical change
  - c. Chemical property
  - d. Physical property
2. A change involving the atomic and molecular structure of a substance
  - a. Chemical change
  - b. Physical change
  - c. Chemical property
  - d. Physical property
3. Property of a compound that can change without involving a change in chemical composition
  - a. Chemical change
  - b. Physical change
  - c. Chemical property
  - d. Physical property
4. Property that changes the chemical make-up of a substance
  - a. Chemical change
  - b. Physical change
  - c. Chemical property
  - d. Physical property

Physical property (P) or Chemical property (C)

5. Melting point \_\_\_\_\_
  6. Flammability \_\_\_\_\_
  7. Boiling point \_\_\_\_\_
  8. Reactivity \_\_\_\_\_
  9. Texture \_\_\_\_\_
  10. Ability to rust \_\_\_\_\_
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11. Substance composed of atoms having an identical number of protons in each nucleus. Cannot be reduced to simpler substances by normal chemical means.
    - a. Pure substance
    - b. Element
    - c. Compound
    - d. Mixture
  12. Substance containing two or more elements and cannot be separated by physical means
    - a. Pure substance

- b. Element
  - c. Compound
  - d. Mixture
13. A sample of matter, either an element or compound, with definite physical properties and a definite composition
- a. Pure substance
  - b. Element
  - c. Compound
  - d. Mixture
14. Composed of two or more substances that are not chemically combined and can be separated.
- a. Pure substance
  - b. Element
  - c. Compound
  - d. Mixture
15. What is a heterogeneous mixture?
- a. Uniform in structure and composition
  - b. Not uniform in structure and composition
16. What is a homogeneous mixture?
- a. Uniform in structure and composition
  - b. Not uniform in structure and composition
17. A homogeneous mixture is often called
- a. Element
  - b. Pure substance
  - c. Solution
18. An alloy is in which mixture group
- a. Homogeneous
  - b. Heterogeneous
19. Give an example of a heterogeneous mixture \_\_\_\_\_
20. Give an example of a homogeneous mixture \_\_\_\_\_

List three means of physically separating the parts of a mixture. Add additional ways for extra credit:

21. \_\_\_\_\_
22. \_\_\_\_\_
23. \_\_\_\_\_
- a. \_\_\_\_\_ (extra credit)
  - b. \_\_\_\_\_ (extra credit)
24. Superheated gas can become
- a. Element
  - b. Plasma
  - c. Compound

d. Mixture

25. What is the melting point? \_\_\_\_\_
26. What is the freezing point? \_\_\_\_\_
27. What is the condensation point? \_\_\_\_\_
28. What is the sublimation point? \_\_\_\_\_
29. Which two of the above are the same? \_\_\_\_\_
30. Briefly describe the Law of the Conservation of Mass

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Complete the following density problems:

31. What is the density of a substance with a mass of 2.0g and a volume of 5 cm<sup>3</sup>?
32. Calculate the mass of a substance with a density of 8 g/cm<sup>3</sup> and a volume of 16 cm<sup>3</sup>?

Extra Credit:

33. What is the density of something that floats in water? \_\_\_\_\_
34. What is the density of something that sinks in water? \_\_\_\_\_
35. What is the density of water? \_\_\_\_\_