

Day 35      Chemistry with Lab      Easy Peasy All-in-One Homeschool

Each orbital can hold a maximum of \_\_\_\_\_ electrons.

Since both electrons have a \_\_\_\_\_ charge, they \_\_\_\_\_.

What keeps them from flying apart? Each electron \_\_\_\_\_ on its axis.

One spins \_\_\_\_\_ and the other spins counter-clockwise. When charged particles

spin, they act like tiny magnets. Since the two electrons spin in \_\_\_\_\_

directions, one acts like the north pole of a magnet and the other acts like the south pole.

This makes the electrons \_\_\_\_\_.

Since each orbital can hold \_\_\_\_\_ electrons:

The "s" sublevel can hold \_\_\_\_\_ electrons.

The "p" sublevel can hold \_\_\_\_\_ electrons.

The "d" sublevel can hold \_\_\_\_\_ electrons.

The "f" sublevel can hold \_\_\_\_\_ electrons.

The orbital shaped like a "dumb-bell" is the \_\_\_\_\_ orbital, while the orbital shaped

spherically is the \_\_\_\_\_ orbital.

How many sublevels are present in the third main energy level? \_\_\_\_\_

What is the maximum number of orbitals in the "d" sublevel? \_\_\_\_\_

The maximum number of electrons that can occupy an orbital is \_\_\_\_\_,

provided they have \_\_\_\_\_ spins.

The highly probable location of an electron within the atom is a(n) \_\_\_\_\_.