

Lesson 35: Electron Distribution Worksheet

Chemistry with Lab

Each orbital can hold a maximum of _____ electrons.

Since both electrons have a _____ charge, they _____.

What keeps them from flying apart? Each electron _____ on its axis.

One spins _____ and the other spins counter-clockwise. When charged particles spin, they act like tiny magnets. Since the two electrons spin in _____ directions, one acts like the north pole of a magnet and the other acts like the south pole.

This makes the electrons _____.

Since each orbital can hold _____ electrons:

The "s" sublevel can hold _____ electrons.

The "p" sublevel can hold _____ electrons.

The "d" sublevel can hold _____ electrons.

The "f" sublevel can hold _____ electrons.

The orbital shaped like a "dumbbell" is the _____ orbital, while the orbital shaped spherically is the _____ orbital.

How many sublevels are present in the third main energy level? _____

What is the maximum number of orbitals in the "d" sublevel? _____

The maximum number of electrons that can occupy an orbital is _____, provided they have _____ spins.

The likely location of an electron within the atom is a(n) _____.