

Lesson 68: Ionic/Covalent ANSWERS

Chemistry with Lab

Determine whether each compound is ionic or covalent. Name the compound.

1. CoBr_2 ionic; cobalt (II) bromide
2. P_2O_5 covalent; diphosphorous pentoxide
3. FeSO_4 ionic; iron (II) sulfate
4. Na_2CO_3 ionic; sodium carbonate
5. P_4 covalent; phosphorous
6. B_2H_4 covalent; diboron tetrahydride
7. CO covalent; carbon monoxide
8. GaCl_3 ionic; gallium chloride
9. SiO_2 covalent; silicon dioxide
10. NH_3 covalent; ammonia

Determine whether each compound is ionic or covalent. Write the formula.

11. dinitrogen trioxide covalent; N_2O_3
12. lithium acetate ionic; $\text{LiC}_2\text{H}_3\text{O}_2$
13. phosphorous trifluoride covalent; PF_3
14. silicon tetrafluoride covalent; SiF_4
15. vanadium (V) oxide ionic; V_2O_5
16. silver phosphate ionic; Ag_3PO_4
17. aluminum hydroxide ionic; $\text{Al}(\text{OH})_3$
18. nitrogen covalent; N_2
19. methane covalent; CH_4
20. zinc sulfide ionic; ZnS