Lesson 102: Limiting Reactants Notes

Stoichiometry Problem Guidelines:

1. _____ the equation

2. convert to _____ of _____ substance

3. MAKE THE _____ USING _____ _____

   FROM BALANCED EQUATION

4. convert to _____

Problem Set One:

Limiting Reactant

_____ used up _____ in a _____ reaction.

Excess Reactant

_____ that is not used up on a chemical _____.

Lesson 102: Limiting Reactants Notes (cont.) Chemistry with Lab

Ex. Problem:

When $FeCl_3$ reacts with O_2, Fe_2O_3 and Cl_2 are produced. If 4.0 moles of $FeCl_3$ and 4.0 moles of O_2 are mixed, how many grams of Fe_2O_3 will be produced?

$$\longrightarrow$$

(Hint: Work two separate problems, using one reactant at a time.)

(Hint: Answer will be the [smaller, larger] amount of product.)

What is the limiting reactant?

What is the excess reactant?

Problem Set Two:

The Chemistry Quiz

CR1. _____ CR2. _____ 1. _____ 2. _____ 3. _____

4. _____ 5. _____

(From GAVL, creative commons 3.0 [https://creativecommons.org/licenses/by/3.0/], https://cms.gavirtualschool.org/Shared/Science/ Chemistry/stoichiometry_shared/Module_8_Limiting_Reactants_Note_Taking_Guide.pdf)