

Lesson 133: Acids, Bases, Salts Notes

Chemistry with Lab

Arrhenius definitions: When reacting with _____,

- _____ release _____ ions (ex _____)
- _____ release _____ ions (ex _____)
- _____ are _____ compounds that release neither _____ ions nor _____ ions

Practice: Write "A" for acid, "B" for base, and "S" for salt.

Solution	Acid (A), Base (B), or Salt (S)	Reacts with Mg (R or NR)	Conductivity (C or NC)
HNO ₃			
NaOH			
NaCl			
H ₂ SO ₄			
HCl			
NH ₄ OH			
KNO ₃			
KOH			
HC ₂ H ₃ O ₂			

Lesson 133: Acids, Bases, Salts Notes (cont.)

Chemistry with Lab

Operational Definitions:

ACIDS:

- _____ taste
- react with some _____ to produce _____ gas
- _____
- turn _____ - _____ indicators different _____
- react with _____ to produce a _____ and water

BASES:

- _____ taste
- _____
- _____
- turn _____ - _____ indicators different _____
- react with _____ to produce a _____ and water

Naming Bases and Salts

To name _____ and _____, follow the standard rules for naming _____ compounds.

NaOH _____

CaCl₂ _____

Lesson 133: Acids, Bases, Salts Notes (cont.)

Chemistry with Lab

Naming Acids

BINARY ACID

- only _____ elements
- first element is _____
- named: hydro_____ic acid
- HCl _____

TERNARY ACID

- _____ elements
- first element is _____
- other elements part of a _____ ion
- naming does NOT require a _____
- _____ \rightarrow _____
- _____ \rightarrow _____
- H_2SO_4 - _____ acid
- H_2SO_3 - _____ acid

Examples:

- H_3PO_4 _____
- HF _____
- HClO_2 _____

Lesson 133: Acids, Bases, Salts Notes (cont.)

Chemistry with Lab

Name the following acids:

_____	_____
_____	_____
_____	_____
_____	_____

The Chemistry Quiz

CR1. _____ CR2. _____ 1. _____ 2. _____ 3. _____

4. _____ 5. _____