

Lesson 172: Le Chatelier's Principle Datasheet

Chemistry with Lab

Table I:

Ion	Color
K ⁺	
Cl ⁻	
SCN ⁻	
Fe ³⁺	
Fe(SCN) ²⁺	

Use the table on the left to determine the colors of reactants and products. Write the color in the blanks above the equation below.

Table II:

colors: _____



Chemical Added	Color Change	Direction of Shift
FeCl ₃		
NaOH		
KSCN		

Conclusions:

According to LeChatelier's Principle, when a _____ is applied to a system in equilibrium, the system will readjust to _____ the stress, restoring a state of equilibrium.

Lesson 172: Le Chatelier's Principle Datasheet (cont.)

Chemistry with Lab

For each procedure in Table II, identify the stress (ex. – addition of a reactant, removal of a product, etc.) and the reason for the shift in equilibrium (ex. – shift to the right uses up reactants):

Stress

Reason for Shift

Hint: NaOH reacts with Fe^{3+} to form solid $\text{Fe}(\text{OH})_3$
